

Hither you will go out all my dead body

I Will Show You Myself by web-cam or We commode encounter!

FIND ME BY NICK: *Dana925*

FIND OUT



$$dE = \delta Q + \delta W$$
 where δQ is the heat supplied to the organisation and δW is the work done on the organisation. Equipartition of energy The muscularity of a mechanical harmonic oscillator (a molecule on a spring) is equally divided between kinetic and potential energy. At any point in the vibration cycle it is exclusively kinetic, and at other points it is exclusively potential. All over the cycle, or all over many cycles, the energy is therefore evenly divided between kinetic and potential energy. This is called the equipartition theorem; the average energy of a system with many degrees of freedom is equally divided among all uncommitted degrees of freedom. This principle is vitally important to understand the behaviour of a measure intimately related to get-up-and-go, called entropy. Entropy is a measure of the invariability of a distribution of activity between the degrees of freedom of a system. When an organisation is tending to a greater extent of freedom (i.e., the number of available get-up-and-go states that are equally accessible), and so the energy is spread completely and evenly among all available degrees of freedom without preference between "one-time" and "one-time" degrees. This numerical outcome is called the Boltzmann constant of thermodynamics. The second basic legal philosophy of thermodynamics is valid only against systems which are good or in equilibrium. Against non-equilibrium systems, the laws governing their behaviour are however disputable. One and only one of the directive principles for these systems is the principle of maximum selective information production.[19][20] It states that non-equilibrium systems behave in such a mode to maximise their randomness yield.